

Quantum theory of interfacial tension quantitatively predicts spontaneous charging of nonpolar aqueous interfaces

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The spontaneous negative charging of aqueous nonpolar interfaces has eluded quantitative first-principle prediction, possibly because it steadfastly challenges the Debye dielectric picture. In this work we show that quantitative prediction requires a substantive revision of Debye's linear dielectric *ansatz* to incorporate an anomalous polarization component yielding electrostatic energy stored as interfacial tension. The minimization of this interfacial tension involves a reduction in hydrogen-bond frustration that takes place upon hydroxide ion adsorption. The quantitative predictions are validated vis-à-vis measurements of hydroxide adsorption uptake obtained using sum-frequency vibrational spectroscopy.

Key Words: interfacial tension, hydrogen bond, hydrogen bond frustration, adsorption, water dielectrics

1. Introduction

Interfacial water on a nonpolar surface remains a subject of intense scrutiny due to its relevance in delineating the molecular forces that steer protein folding [1] and molecular associations [2], and drive biotechnological innovation [2-6]. The spontaneous negative charging of such interfaces through transference of hydroxide ions (OH^-) from bulk water has been established [7-9], yet the topic still remains a subject of contention [5, 7, 9]. No first-principle inference of spontaneous interfacial charging has so far generated quantitative predictions amenable of experimental verification [8]. In this regard, the treatment of interfacial water through molecular dynamics (MD) appears to yield contradictory results, even predictions of acidic (proton donating) interfaces [7, 10]. This controversy stimulates the fundamental approach taken in this work, which involves a substantive revision of Debye's dielectric ansatz [4]. The latter would predict no polarization-induced charge since there is no intrinsic electrostatic field to speak of. Thus, we may state that the spontaneous charging of an aqueous nonpolar interface poses a challenge to the standard dielectric picture, demanding a reworking of the Debye ansatz starting from first principles.

In this work we argue that experimentally verifiable quantitative predictions of spontaneous negative charging of aqueous hydrophobic interfaces require a general treatment of interfacial water dielectrics that cannot be subject to the constraints imposed by Debye's ansatz regarding polarization alignment [1,11]. The conceptual departure is essential to include the anomalous polarization component that arises as a result of hydrogen-bond frustration at the contact region between the two bulk phases. This revision could not have been reasonably envisioned at the time Debye's linear dielectric ansatz was formulated since details on water hydrogen-bonding

structure were unknown [4]. As shown subsequently, the non-Debye contribution generates interfacial tension that is stored as electrostatic energy, a tension shown in this work to be mitigated by the adsorption of hydroxide ions. By frustration in this context we refer to a hydrogen bonding opportunity that cannot materialize because the proton acceptor or nucleophilic group is not physically present, as it is the case with the water layer in contact with the nonpolar surface.

As previously noted, the departure from the Debye linear dielectric *ansatz* becomes imperative because no charge $\gamma = -\nabla \cdot \mathbf{P}$ may be generated by water polarization (\mathbf{P}) if the latter is made proportional to the vanishing electrostatic field \mathbf{E} at the nonpolar surface, as it is the case with the Debye assumption $\mathbf{P} = (\epsilon - \epsilon_0)\mathbf{E}$, where ϵ , ϵ_0 denote dielectric and vacuum permittivity, respectively.

To predict quantitatively the spontaneous charging of the aqueous nonpolar interface, a more general dielectric scheme is required. The hydrogen-bond frustration of interfacial water, described by the scalar field $\phi = \phi(\mathbf{r})$, with $\phi(\mathbf{r})$ = expected number of unfulfilled hydrogen bonds for a water molecule at position \mathbf{r} , generates a non-Debye polarization component $\mathbf{P}^\# = \mathbf{P} - (\mathbf{P} \cdot \mathbf{e})\mathbf{e}$ ($\mathbf{e} = \mathbf{E}/\|\mathbf{E}\|$). In turn, this “orthogonal” polarization induces a net charge $\gamma^\# = -\nabla \cdot \mathbf{P}^\#$. In previous work [1,11] we showed that distortions from bulk structure cause $\mathbf{P}^\#$ to be proportional to the frustration gradient, according to the equilibrium relation

$$\mathbf{P}^\# = -\xi \nabla \phi, \quad (1)$$

where $\xi = (\lambda \epsilon_0)^{1/2}$ and $\lambda = 9.0 \text{ mJ/m}$ at 298K [1]. The generation of net charge, and thereby of polarization, arises from the hydrogen-bond frustration at the interface (Eq. 1) that leaves partial charges on the water atoms untitrated. The lack of interfacial charge neutralization implies that interfacial tension, which by definition arises from hydrogen-bond frustration, is also stored

electrostatically in the anomalous polarization [1, 4], as mathematically shown in this work. Eq. 1 will be subsequently corroborated in the context of this study. From Eq. 1, the net charge $\gamma^\#$ induced by $\mathbf{P}^\#$ becomes $\gamma^\# = \xi \nabla^2 \phi$. By making interfacial tension proportional to electrostatic energy, the non-Debye treatment enables the computation of the net charge induced by the frustration at the water layer in contact with the nonpolar surface, and hence becomes adequate to predict quantitatively the spontaneous charging of the interface.

2. Methods

For simplicity, we consider water in contact with a featureless slab on nonpolar nonpolarizable material typically realized by hydrogenated graphene [9]. The frustration or net loss of hydrogen bonding opportunities at the hydrophobic interface generates interfacial tension quantifying the disruption of the bulk-like tetrahedral hydrogen-bond coordination of water [1, 2, 11]. The extent of frustration of a water molecule at a liquid/solid interface is expected to be higher than that of a hydroxide ion (OH^-) since the donated proton in the latter chemical species is expected to have a lower electrophilicity than the water proton, and therefore is expected to be less prone to form hydrogen bonds with a nearby nucleophile [9]. This conjecture will be corroborated in this work and is based on the fact that the hydroxide oxygen is less electrophilic than the water oxygen, since the former gained an electron while the latter shares an electron pair with a hydrogen atom. The argument leads us to postulate that the interfacial tension is lowered by adsorption of hydroxide ions or, in other words, the spontaneous negative charging of the nonpolar interface is expected to result from the relief of interfacial tension.

Given the above-noted observations, a quantum parameter, ζ , will be subsumed into the MD computations scaling a term accounting for interfacial energy in order to quantify the degree of

frustration of interfacial hydroxide relative to interfacial water. The relative hydrogen bond frustration is given by the quotient $\zeta = \zeta(\text{OH}^-/\text{H}_2\text{O}) = F_h/F_w = 0.021/0.179 = 0.117 (\pm 0.017)$, where F_h , F_w are the expected net quantum electrostatic fields (in $\text{e}/\text{\AA}^2$ units) at the position or probe site of a putative hydrogen-bonding nucleophile (lone electron pair) that acts as putative acceptor of the proton donated by OH^- and H_2O , respectively [9]. The fields were obtained by molecular projection of the delocalized quantum charges in the aqueous condensed phase [12] along the direction of proton donation for hydrogen bonding [9]. The authors in ref. 9 implemented a Car-Parrinello molecular dynamics scheme in a Wannier representation of the condensed phase [12], thus generating a unique set of maximally localized Wannier functions that realized the solid-state equivalent of localized molecular orbitals.

Based on the estimation of the quantum parameter ζ given above, we may assert that *the extent of frustration due to an unfulfilled hydrogen bond in the hydroxide ion at a nonpolar interface is 11.7% of the hydrogen-bond frustration of water with a dangling OH group*. Conversely, $\zeta(\text{H}_3\text{O}^+/\text{H}_2\text{O}) = 1.922 (\pm 0.102)$, (cf. [9]) implying that hydronium adsorption would entail almost twice as much frustration than water at the interface. To summarize, mitigation of interfacial tension dictates that hydroxide ion - and not hydronium - must be preferentially adsorbed.

3. Theory/calculations

To predict the spontaneous charging of the hydrophobic interface, we first incorporate into the MD simulations the interfacial energy per unit volume, Δu_ζ , associated with spanning a water interface that envelops a solid non-polarizable hydrophobic slab of macroscopic dimensions. The energy-density term Δu_ζ quantifies the distortion of the water hydrogen-bond matrix and is therefore given by the elastic integral [1]:

$$\Delta u_\zeta = (1/2)\lambda \int \|\nabla\phi\|^2 d\mathbf{r}, \quad (2)$$

where $\phi=\phi(\mathbf{r})$ describes the local time-averaged hydrogen-bond frustration of the solvent and the integration is carried over a spatial domain Ω large enough so that its border $\partial\Omega$ is fully contained in bulk water, where the following conditions are satisfied: $\phi(\mathbf{r})=0$, $\nabla\phi(\mathbf{r})=0 \ \forall \mathbf{r} \in \partial\Omega$. To describe the local distortion of water structure at spatial location \mathbf{r} relative to the bulk hydrogen-bond pattern, we introduce the scalar field $\phi=\phi(\mathbf{r})$ ($0 \leq \phi \leq 4$) rigorously representing the expected frustration of a water molecule or a hydroxide ion while any of them visits a sphere of radius $r=4\text{\AA}$ centered at position \mathbf{r} for a timespan $\tau=1\text{ps}$ [1, 2,11]. If the proton donation of the hydroxide ion leading to hydrogen bonding is frustrated due the absence of a proton acceptor, we get $\phi=\zeta$ (<1) while the frustration of water hydrogen bonding yields $\phi=1$.

In essence, we compute the interfacial tension arising from the hydrogen bonding frustration of interfacial water, and take into account the fact that this tension is minimized by the adsorbed hydroxide ions, whose dangling donated proton is 88.3% less frustrated than the water counterpart (Fig. 1). Our theoretical treatment may be benchmarked and validated, as done in this work, since progress in spectroscopic methods based on second-order nonlinear optics enable direct examination of water structure at the interface [8].

The quantum-mechanical parameter ζ is thus incorporated in the MD simulations to account for the reduction in the frustration levels on the first layer directly in contact with the nonpolar surface due to hydroxide ion adsorption.

To validate Eq. 1, we observe that this equation dictates that the interfacial tension is actually stored at equilibrium as the non-Debye polarization energy density since

$$\Delta u_\zeta = (1/2)\lambda \int \|\nabla\phi\|^2 d\mathbf{r} = (1/2)\epsilon_0^{-1} \int \mathbf{P}^{\#2} d\mathbf{r} \quad (3)$$

The direct MD computation of the time evolution of the electrostatic energy density and interfacial tension has been performed by incorporating the global term Δu_ζ in the potential energy stepwise evaluation. The convergence of the two quantities has been observed (Fig. 2), thus validating the equilibrium relation presented in Eq. 1.

Since the surface integral $\oint [\phi \nabla \phi] \cdot d\sigma$ over $\partial\Omega$ vanishes as per the choice of Ω , integration of Eq. 2 by parts yields

$$0 \leq \Delta u_\zeta = -(1/2)\lambda \int \phi \nabla^2 \phi d\mathbf{r} = -(1/2)(\lambda/\epsilon_0)^{1/2} \int \gamma^\# \phi d\mathbf{r} \quad (4)$$

Since $\phi \geq 0$, we obtain:

$$\gamma^\#_{MV} = -(\lambda\epsilon_0)^{1/2} \int \|\nabla \phi\|^2 d\mathbf{r} / \int \phi d\mathbf{r} \leq 0, \quad (5)$$

where $\gamma^\#_{MV}$ is the mean-value induced charge per unit volume yielding: $\gamma^\#_{MV} \int \phi d\mathbf{r} = \int \gamma^\# \phi d\mathbf{r}$.

We have then proven the following *theorem*:

An aqueous featureless nonpolar interface causes hydrogen-bond frustration and generates a negative polarization-induced charge, yielding $\gamma^\#_{MV} \leq 0$.

4. Results

The net induced charge at the interface results from the adsorption of the negatively charged hydroxide ions. The net charge was estimated according to Eq. 5, requiring the construction of the field $\{\phi(\mathbf{r}), \nabla \phi(\mathbf{r})\}$ (cf. ref. 1, 11, Eq. 5) within an NPT-statistical ensemble for a featureless hydrophobic interface in the limit of infinite curvature. Thus, the calculation of $\gamma^\#_{MV}$ following Eq. 5 yields $[\text{OH}^-]^\# = \gamma^\#_{MV}/Y = 15.4\text{mM}$, where $[\text{OH}^-]^\#$ = hydroxyl molar concentration at interface and Y is the net charge per mole of hydroxyl ion. Thus, at pH 7 (neutral solution) we get the equilibrium constant

$$K=[OH^-]^\#/[OH^-] = 153,832 \quad (6)$$

Let $M=[H_2O] \gg [OH^-]$ be the molar concentration of “free adsorption sites” ≈ 55.5 mole/liter at $T=298K$, then $K=M^{-1}\exp(-\Delta G/RT)$, where ΔG is the free energy change associated with the transference of a hydroxide ion from bulk water to the interface. Thus, our predicted value $\Delta G = -40.02kJ/mole$ compares satisfactorily within experimental error with the experimental value $\Delta G = -45kJ/mole$ obtained from the Langmuir adsorption isotherm monitored by phase-sensitive sum-frequency vibrational spectroscopy [8].

5. Discussion

The spontaneous negative charging of the aqueous interface sustained at a flat nonpolar surface is a fact that challenges and hence motivates a substantive revision of Debye’s dielectric picture, whereby water polarization aligns with the intrinsic electrostatic field. This is so simply because the latter vanishes in the case under study. Thus, the hydrogen bonding frustration of water at a nonpolar interface generates a non-Debye polarization component that stores electrostatic energy as interfacial tension. As shown in this work, the reduction of the interfacial tension is caused by a net adsorption of hydroxide ions resulting in a spontaneous negative charging of the interface. The adsorption of the hydroxide ion is favored on quantum mechanical grounds as its hydrogen bond frustration at the interface is 88.3% lower than that of water. The quantum mechanical basis for the induced interfacial effect within a non-Debye dielectric picture is quantitatively validated vis-à-vis surface measurements using second-order nonlinear optics.

6. Conclusions

The spontaneous negative charging of aqueous nonpolar interfaces has eluded quantitative first-principle prediction because this phenomenon challenges the standard Debye dielectric picture. We showed that quantitative prediction required a reworking of Debye's linear dielectric *ansatz* to incorporate an anomalous polarization component yielding electrostatic energy stored as interfacial tension. To gain interfacial stability, the minimization of interfacial tension becomes operative promoting a reduction in hydrogen-bond frustration. The latter takes place upon hydroxide ion adsorption. The quantitative predictions were validated vis-à-vis experimental measurements of hydroxide adsorption uptake.

Figure Caption

Figure 1. Scheme of interfacial tension quenching due to hydroxide ion adsorption onto a nonpolar aqueous interface. The interfacial region is represented with color gradient, bulk water is in solid blue and the nonpolar slab is in brown. The frustrated hydrogen-bonding opportunity is symbolized by a crossed dotted line. The level of frustration at the interface with the nonpolar slab is reduced by a factor ζ (<1) upon hydroxide ion adsorption, while the interfacial energy density change associated with transference from the bulk to the interface is reduced by a factor ζ^2 due to the square dependence on the frustration gradient of the elastic integrand in Eq. 2.

Figure 2. Convergence of the interfacial tension (Eq. 2) and non-Debye electrostatic energy density (given by the expression $(1/2)\epsilon_0^{-1} \int \mathbf{P}^{\#2} d\mathbf{r}$) as equilibrium is reached. The curves displayed are averages over 10 MD 24ps-trajectories for water dynamics within an isobaric/isothermal ensemble (1atm, 298K) with 10,160 water molecules and a total of 160 ions,

with 80 OH^- and 80 H^+ ($[\text{OH}^-]=0.4\text{M}$) packed against a featureless layer of hydrogenated graphene. The MD 24ps-trajectories for water dynamics are generated as described in ref. 1 except for the incorporation in the potential energy of an interfacial term to account for hydrogen bond frustration at the nonpolar region of contact. The solvent bath provides at least four water layers of solvent envelope with water molecules extending 18\AA from the graphene surface. The evaluation of the interfacial tension-generating frustration incorporates a quantum mechanical parameter ζ describing the relative electrophilicity of the donated hydroxide proton relative to the water proton, a measure of the hydrogen bonding propensity of the dangling O-H bond in hydroxide relative to water.

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FIGURE 1

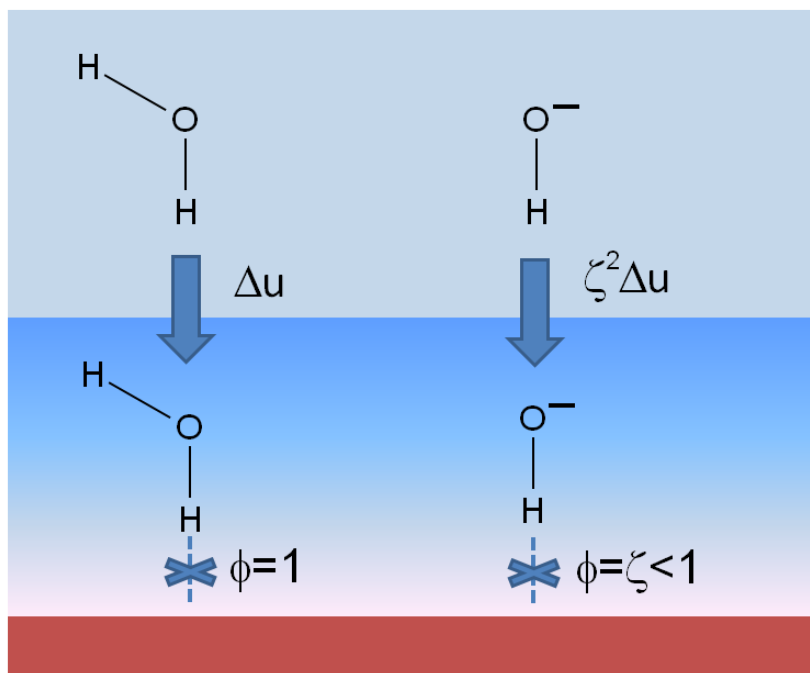


FIGURE 2

